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The Dependence of Tropospheric Ozone Production Rate on Ozone Precursors

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Abstract

An analytic formula is derived expressing the tropospheric O_3 production rate, $P(O_3)$, as a power law function of radical production rate, NO_x concentration, and VOC – OH reactivity. Power law exponents depend on a single parameter, L_N/Q, which is the fraction of free radicals removed by reactions with NO_x. The formula reproduces the functional form of $P(O_3)$ obtained from photochemical box model calculations. Ozone production rates are shown to have a smooth transition between previously derived low and high NO_x limits. Potential applications of this formula include analysis of day to day and place to place variations in $P(O_3)$, with $P(O_3)$ either obtained from measurements collected during field campaigns or produced as output from chemical-transport models.

1. Introduction

A by now standard exercise in most comprehensive photochemistry field campaigns is to use observed chemical concentrations to calculate a chemical rate of O_3 production, $P(O_3)$. There are 3 ways of proceeding. $P(O_3)$ can be determined as the rate of reaction of peroxy radicals with NO based on measured NO and peroxy radicals [e.g. *Cantrell et al*., 1996; *Penkett et al.*, 1999; *Mihelcic et al.*, 2003], $P(O_3)$ can be calculated from a photochemical model that is constrained by observed concentrations of non-radical species [e.g. *Frost et al.*, 1998; *Carslaw et al.*, 2001; *Kleinman et al.*, 2002a], and $P(O_3)$ can be determined from the photostationary state relations [e.g. *Ridley et al.*, 1992; *Hauglustaine et al.*, 1996; *Volz-Thomas et al.*, 2003], albeit with problems. The comparison of $P(O_3)$'s calculated in these ways is an active area of research and provides a stringent test of our understanding of tropospheric O_3 photochemistry. $P(O_3)$ is also of interest for determining the chemical component of O_3 tendency in different regions [e.g. *Liu et al.*, 1992; *Jacob et al.*, 1996; *Mauzerall et al.*, 1996; *Cantrell et al.*, 2003]. The dependence of $P(O_3)$ on O_3 precursors provides information of the O_3 forming process in regions where O_3 is a pollutant with health and other environmental impacts [e.g. *Daum et al.*, 2000a,b, 2003; *Dommen et al.*, 2002; *Thornton et al.*, 2002; *Kleinman et al.*, 2002a]. *Tonnesen and Dennis* [2000a] have advocated that emission control strategies should be targeted towards reducing $P(O_3)$ in regions where it is high.

In order to understand why $P(O_3)$ varies from time to time and place to place it is desirable to have simple, transparent formulas that show how $P(O_3)$ depends on its precursors. For the limiting cases in which NO_x concentration is very high or very low such formulas already exist and have been used to interpret data from field campaigns [e.g. *Kleinman et al.*, 1994; *Carpenter et al.*, 1997, 2000; *Zanis et al.*, 1999; *Daum et al.*, 2000a,b; *Dommen et al.*, 2002]. Formulas valid for mid-NO_x conditions have been derived for the simplified case of a CO/CH_4 atmosphere [*Carpenter et al.*, 2000]. In this article we derive a general relation between $P(O_3)$ and the concentration of the

atmospheric reactants responsible for its production. These reactants can be divided into 3 somewhat overlapping categories, NO_x, VOCs, and radical precursors. Our result will be a formula which gives the power law dependence of $P(O_3)$ on NO_x concentration, VOC reactivity, and radical production rate. This formula depends on a single quantity, L_N/Q, which is the fraction of free radicals removed from the atmosphere by reaction with NO_x, the remaining fraction being removed by combination reactions between free radicals.

While NO_x concentration, VOC reactivity, and radical production rate are quantities that can be directly related to commonly made observations, L_N/Q is a parameter with less intuitive appeal. It is clearly not an independent precursor variable in the same sense as NO_x concentration, VOC reactivity, and radical production rate; as it can be described in terms of those 3 variables [*Kleinman et al.*, 2001]. Instead it is a dependent variable that conveys information on where an air parcel is in the continuum of states ranging from the high NO_x regime characteristic of polluted conditions to the low NO_x regime characteristic of rural or remote regions. Under normal circumstances a polluted air mass evolves from a high NO_x state with L_N/Q near 1 to a low NO_x state with L_N/Q near 0 due to chemical oxidation of VOCs and NO_x and due to dilution with cleaner background air. It is know that the evolution of an airmass from a high NO_x state to a low NO_x state is accompanied by a transition between VOC sensitive O_3 chemistry to NO_x sensitive O_3 chemistry. Because the dominant end product of photochemistry changes from NO_x oxidation species such as HNO_3 to radical combination species such as peroxides during this transition, the ratio of these 2 compounds can be used as an Indicator of NO_x and VOC limited photochemistry [*Sillman*, 1995; *Tonnesen and Dennis*, 2000b; *Hammer et al.*, 2002; *Martilli et al.*, 2002; *Sillman and He*, 2002; *Thielmann et al.*, 2002].

Our objective in this article is to present and evaluate a general formula that shows how $P(O_3)$ depends on its precursors. This work builds on an earlier study [*Kleinman et al.*, 1997] in which the sensitivity of $P(O_3)$ to NO_x concentration and VOC reactivity was derived. Here we add an additional equation that shows the sensitivity of $P(O_3)$ to radical production rate. It is then straight-forward to obtain a formula that shows the power law dependence of $P(O_3)$ on NO_x, VOC reactivity, and radical production rate. Our discussion of the $P(O_3)$ formula focuses on the governing photochemical reactions and the approximations made in the derivation including the assumption that PAN is in steady state.

Evaluation of the $P(O_3)$ formula is done using a combination of observed and predicted quantities. NO_x concentration, VOC reactivity, and radical production rate are determined from observed concentrations, solar irradiance, and published rate constants. A box model with species concentrations constrained to observed values [*Kleinman et al.*, 2002a,b] is used to calculate $P(O_3)$ and L_N/Q. A test of the approximate power law formula is a determination of whether or not calculated $P(O_3)$ has the predicted functional form.

The power law formula succeeds in quantitatively reproducing most of the model results. However, we offer this formula only as a qualitative tool to understand $P(O_3)$ variations as our analysis does not address questions of model completeness or accuracy in the same way that, for example, a closure experiment would.

2. Theory

Qualitative formulas have been derived elsewhere for the limiting low and high NO_x cases [*Kleinman et al.*, 1994; *Daum et al.*, 2000a,b]. In this section we extend the derivation to show the dependence of $P(O_3)$ on radical production rate, NO_x concentration, and VOC reactivity for cases intermediate between the low and high NO_x limits.

2.1 $P(O_3)$ Formula

Derivations are based on the usual photochemical equations [e.g. *Seinfeld and Pandis*, 1997], a subset of which are repeated below. Equations are divided into 3 categories representing the initiation, cycling, and termination steps in the overall O_3 forming chain reaction.

Initiation

R1 $O_3 + h\nu \rightarrow O^1D$

$O^1D + H_2O \rightarrow 2\ OH$

R2 $HCHO + h\nu \rightarrow 2\ HO_2$

R3 $Alkene + O_3 \rightarrow radicals$

Chain Propagation or Cycling

R4 $OH + VOC_i \rightarrow \alpha\ HO_2 + \beta\ RO_2 + \gamma\ RCO_3$

R5 $HO_2 + NO \rightarrow NO_2 + OH$

R6 $RO_2 + NO \rightarrow NO_2 + HO_2 +$ HC oxidation product

R7 $RCO_3 + NO \rightarrow NO_2 + RO_2 +$ HC oxidation product

R8 $NO_2 + h\nu \rightarrow NO + O$

$O + O_2 \rightarrow O_3$

Termination

radical + NO_x

R9 $OH + NO_2 \rightarrow HNO_3$

R10 $RO_2 + NO \rightarrow$ organic nitrate

R11a $CH_3CO_3 + NO_2 \rightarrow PAN$

R11b $PAN \rightarrow CH_3CO_3 + NO_2$

radical + radical

R12 $HO_2 + HO_2 \rightarrow H_2O_2$

R13 $RO_2 + HO_2 \rightarrow ROOH$

R14 $OH + HO_2 \rightarrow H_2O + O_2$

R15 $RO_2 + R'O_2 \rightarrow$ products

Photolysis and ozoneolysis reactions split chemical bonds, forming free radicals that initiate photochemical O_3 production. R1 and R2 in that order are usually the most important initiation reaction. The sum total radical production rate will be denoted as Q.

Chain propagation is a cyclic pathway in which OH oxidizes VOCs (R4) forming peroxy radicals which convert NO to NO_2 (R5-R7) which is photolyzed forming O_3 (R8). The cycle is completed by regenerating OH in R5 and NO in R8. In R4, VOC_i could be a hydrocarbon, CO, an oxygenated compound such as HCHO, or any other species that reacts with OH producing peroxy radicals. The OH reactivity of a mixture of VOCs is given by,

$$VOC_R = \sum k_i [VOC_i] \quad (1)$$

where the k's are rate constants for reaction with OH. Although VOC reactivity and radical production are treated as independent variables in the production of O_3, these 2 precursor categories overlap as compounds such as HCHO contribute to both. R4 is written to allow for the production of 3 types of peroxy radicals, not all of which would be produced from any single VOC. The total number of NO to NO_2 conversions caused by the peroxy radicals formed in R4 assuming that they all react with NO as in R5-R7 is given by

$$Y_i = \alpha + 2\beta + 3\gamma \quad (2)$$

Ozone production stops with the chain termination reactions in which radicals are destroyed. Conservation of radicals requires that the rate of termination equal the rate of initiation. There are two families of termination reactions in which free radicals either

react with NO_x or with each other. An important parameter describing O_3 photochemistry is the fraction of radicals removed in each pathway. For this purpose we use L_N/Q where L_N is the rate of removal by reactions of the radical + NO_x type. Formation of PAN removes radicals in R11a while PAN dissociation is a radical source in R11b. We will use the symbol P(PAN) to indicate the net PAN production rate given as the rate of R11a minus R11b. Otherwise, the symbol "P" will stand for a chemical production rate and not include loss processes.

The derivation of formulas for $P(O_3)$ depends on several assumptions. In the low NO_x case defined by $L_N/Q \rightarrow 0$, the rate limiting step is taken to be reaction of peroxy radical with NO (R5-R7). It is assumed that radical loss by peroxide formation is equal to radical production, yielding

$$P(O_3) = k_t/(2\ k_{eff})^{1/2}\ Q^{1/2}\ [NO] \qquad (3)$$

where k_t is a composite rate constant for R5-R7 and k_{eff} describes the effective rate of peroxide formation [*Kleinman et al.*, 1994]. Under high NO_x conditions, $L_N/Q \rightarrow 1$ and $P(O_3)$ is given by

$$P(O_3) = Q \sum Y_i\ k_i\ [VOC_i] / k_9\ [NO_2] \qquad (4)$$

[*Daum et al.*, 2000a,b]. The major assumptions used to derive (4) are that the reaction of OH with VOCs (R4) are rate limiting [Sillman, 1995],

$$P(O_3) \approx \sum Y_i\ k_i\ [OH]\ [VOC_i] \qquad (5)$$

that radical + radical reactions (R12-R15) can be ignored, and that the reaction of OH with NO_2 (R9) is the predominant radical sink:

$$L_N \approx P(HNO_3) = k_9\ [OH]\ [NO_2] \qquad (6)$$

A general formula for $P(O_3)$ valid at intermediate NO_x concentration is based on equations for the relative sensitivity of $P(O_3)$ to its precursors. The sensitivity to NO_x and VOCs is given in *Kleinman et al.* [1997]:

$$d\ln P(O_3)/d\ln[NO_x] = (1 - 3/2\ L_N/Q) / (1 - 1/2\ L_N/Q) \quad (7)$$

$$d\ln P(O_3)/d\ln(VOC_R) = 1/2\ L_N/Q / (1 - 1/2\ L_N/Q) \quad (8)$$

Equations (7) and (8) express sensitivities in terms of logarithmic derivatives, i.e. $d\ln P/d\ln X = (X/P)dP/dX$, with the property that $d\ln P/d\ln X=1$ when a (small) n% change in X produces an n% change in P. Equation (7) is written with NO_x as the independent variable rather than NO as originally derived. In deriving (7) we had assumed that NO and NO_2 are proportional, an approximation that is discussed in a following section. In the Appendix we derive an analogous expression for the sensitivity of $P(O_3)$ to radical production rate:

$$d\ln P(O_3)/d\ln Q = 1/2 / (1 - 1/2\ L_N/Q) \quad (9)$$

The derivation of (9) in the Appendix closely follows that used to obtain (7) and (8) and serves to illustrate the approximations made in arriving at those equations.

We obtain our desired general formula by writing $P(O_3)$ as:

$$P(O_3) = K\ Q^{C1}\ [NO_x]^{C2}\ (VOC_R)^{C3} \quad (10)$$

where K has zero first derivatives but varies with L_N/Q. Differentiating (10) with respect to Q, NO_x or VOCs yields, respectively

$$d\ln P(O_3)/d\ln Q = C1 \quad (11)$$

$$d\ln P(O_3)/d\ln[NO_x] = C2 \quad (12)$$

$$d\ln P(O_3)/d\ln(VOC_R) = C3 \quad (13)$$

The exponents C1, C2, and C3 depend only on L_N/Q and are given by (7-9). A graph of these 3 functions is shown in Fig. 1. As indicated in Table 1, C1, C2, and C3 have the limiting low and high NO_x values given in (3) and (4), respectively. At $L_N/Q = 1/2$, the NO_x and VOC curves cross indicating an equal sensitivity of $P(O_3)$ to both precursors. The NO_x sensitivity curve has a zero crossing at $L_N/Q = 2/3$. At this point $P(O_3)$ has a maximum value with respect to changes in NO_x concentration. An analogous feature appearing in ozone isopleth plots is the O_3 ridge line [e.g. *Seinfeld and Pandis*, Fig. 5.15, 1997].

2.2 NO or NO_2?

In the derivation of (7-9) it is assumed that NO and NO_2 are proportional, an approximation which can be improved upon. NO_x enters the derivation of the sensitivity equations in 2 places. In (A1), $P(O_3)$ is given by the rate of reaction of peroxy radicals with NO and in (A7) L_N is given by the rate of reaction of OH with NO_2. The limiting low and high NO_x formulas for $P(O_3)$, (3) and (4), show that this dual dependence on NO and NO_2 reduces to a dependence on NO at low NO_x and a dependence on NO_2 at high NO_x. This feature can be incorporated in (10) by using NO as an independent variable at low L_N/Q and NO_2 at high L_N/Q. In a following section we empirically determine the values of L_N/Q which give better results with NO or NO_2 as an independent variable.

2.3 PAN

PAN and analogous compounds are temporary reservoirs for NO_x. Formation of reservoir compounds removes NO_x and free radicals in one location, almost all of which is put back into the atmosphere at another location when theses compounds dissociate. As discussed in *Kleinman et al.* [2002b], PAN and analogous compounds have not been included in our calculations on the assumption that these species are in steady state.

Under steady state conditions the rate of formation of PAN is equal to the rate of dissociation and PAN is then neither a source or sink of free radicals. Although we expect this to be true on average, most of our studies have been done close to emission sources in air masses that have an active photochemistry, where we expect PAN to be a net radical sink. Our ability to improve upon the steady state PAN approximation has been limited by the absence of PAN measurements in our data sets.

The effect of PAN on $P(O_3)$ sensitivities has been addressed by *Sillman* [1995], *Spirig et al.* [2002] and *Sillman and He* [2002] and indirectly by *Kleinman et al.* [1997]. *Sillman* [1995] noted that the point at which $P(O_3)$ has an equal sensitivity to NO_x and VOC occurs when

$$P(HNO_3) = 2\ P(\text{Peroxide}) \qquad (14)$$

i.e., where the number of radicals removed by forming HNO_3 is equal to the number removed by forming peroxides. Figure 1 and Table 1, derived under the assumption that PAN is in steady state, indicate that equal sensitivity to NO_x and VOC occurs at $L_N/Q = 1/2$, a result that reduces to Sillman's formula when formation of organic nitrates and PAN is ignored (when $L_N = P(HNO_3)$). In deriving (14), Sillman assumed that PAN and O_3 production rates are proportional. Using the same approximation, equations in *Kleinman et al.* [1997] show that the sensitivity of $P(O_3)$ to NO_x and VOCs is given by

$$d\ln P(O_3)/d\ln[NO_x] = (1 - 3/2\ L_N/Q + 1/2\ P(PAN)/Q) / (1 - 1/2\ L_N/Q) \qquad (15)$$

$$d\ln P(O_3)/d\ln(VOC_R) = (1/2\ L_N/Q - 1/2\ P(PAN)/Q)/ (1 - 1/2\ L_N/Q) \qquad (16)$$

Equating the right hand sides of (15) and (16) to determine the point at which NO_x and VOC sensitivity are equal yields (14) after a change in notation. Applying the same approximation to $d\ln P(O_3)/d\ln Q$ does not change (9). Equation (10) still applies except that C2 and C3 are given by the right hand sides of (15) and (16) instead of (7) and (8).

For positive values of P(PAN), (15) and (16) show that NO_x sensitivity is increased and VOC sensitivity reduced, relative to the no – PAN case. *Spirig et al.* [2002] have a numerical example of this effect in a case with high PAN concentrations (PAN ≈ HNO_3 ≈ 10 ppb). They find that a calculation without PAN overpredicts VOC sensitivity, relative to a calculation that includes PAN, in regions where PAN is forming, and overpredicts NO_x sensitivity in regions where PAN is dissociating. A comparison of Figs. 16a and b in *Spirig et al.*, shows these over and underpredictions cause a 1 hour shift in a transition between VOC and NO_x sensitive photochemistry but do not change the qualitative sense of the calculations.

2.4 L_N/Q

According to (7) – (10) the functional form of $P(O_3)$ is determined entirely by L_N/Q. In this section we show how L_N/Q depends on O_3 precursors. We rely on formulas given in *Kleinman et al.* [2001]. Using the same approximations that were employed in deriving the $P(O_3)$ sensitivity equations (see Appendix), L_N/Q is given by the solution of a quadratic equation:

$$L_N/Q = -\alpha/2 + (\alpha^2 + 4\,\alpha)^{1/2} / 2 \qquad (17)$$

where

$$\alpha = (k_9\,[NO_2]k_5\,\gamma'\,[NO]/(VOC_R))^2\,(1/2\,Q\,k_{eff}) \qquad (18)$$

$$\gamma' = [HO_2]/([HO_2]+[RO_2]) \qquad (19)$$

In a following section we will show the relation between L_N/Q and NO_x for 2 cities, Phoenix and Houston. Here we note a few qualitative feature. Equations (17) - (19) depend on the HO_2 to total peroxy radical ratio through γ' and k_{eff}. However, that dependence is weak and a reasonable guess for the split between HO_2 and RO_2 will

usually suffice. If we assume that NO and NO_2 are proportional, then L_N/Q is a monotonic function of $NO_x^4/(VOC_R)^2$ (1/Q). Because NO_x appears to the fourth power in this expression, NO_x concentration more than any other variable determines L_N/Q.

3. Model

A constrained steady state (CSS) box model is used to calculate $P(O_3)$ and L_N/Q. This is the same model used in previous studies in which aircraft observations in Nashville, New York City, Phoenix, Philadelphia, and Houston were used to determine O_3 production rates in those cities [*Kleinman et al.*, 2002a]. Inputs to the CSS model include the chemical species; O_3, NO, CO, speciated hydrocarbons, HCHO, H_2O_2, organic peroxides, SO_2, and water vapor; actinic flux approximated from a UV measuring Eppley radiometer [*Madronich*,1987] ; temperature; and pressure. PAN and analogous compounds are assumed to be in steady state and are not included as calculated or constrained variables. Calculations are based on the RADM2 chemical mechanism [*Stockwell et al.*, 1990] for anthropogenic pollutants and the mechanism of *Paulson and Seinfeld* [1992] for isoprene oxidation. Observed concentrations of approximately 100 VOCs are parsed into RADM2 categories according to structure and reactivity. OH reaction rate constants are taken from *Atkinson* [1994] where available, else estimated from structure reactivity relations according to the methods of *Kwok and Atkinson* [1995].

The CSS calculations yield the concentrations of OH, HO_2, RO_2s and NO_2 which are in steady state with the observed mixture of stable atmospheric species. The model keeps track of the rates of all individual reactions. $P(O_3)$ is determined as the rate for conversion of NO to NO_2 by peroxy radicals which is the sum of R5, R6, and R7. L_N is the rate of radical loss due to R9 and R10, Q is the rate of radical production due to all photolysis reaction (2 of which are R1 and R2) plus the rate due to ozoneolysis reactions, R3.

4. Experimental

Chemical concentrations and other input parameters for the CSS calculations are taken from aircraft observations made during the course of 5 field campaigns in Nashville, New York City, Phoenix, Philadelphia, and Houston. VOC observations determined by canister samples are generally the limiting factor in assembling the input data to a CSS calculations. For each VOC sample, other trace gasses and parameters are averaged over the sample time period (about 30s). Experimental techniques and data sets have been previously described ([*Kleinman et al.*, 2002a,b; *Daum et al.*, 2003] and references therein).

A few general features of the 5 city data set should be noted. Samples used in this study are all taken during daylight hours. Most samples are at mid-boundary layer height. About 5% are in much cleaner air above the boundary layer. There is a wide variation in chemical concentrations and physical conditions, some of which has been explored in previous articles and some of which will be the topic of a future study comparing O_3 production in the 5 cities. Among the distinctive features are very high VOC reactivity in Houston, a very dry atmosphere and low radical production rate in Phoenix, high isoprene concentrations in Nashville, and high NO_x concentrations in power plant plumes. Although our objective in this study is not to describe the 5 cities, it is important to note that the conditions under which sampling was done and $P(O_3)$ calculated are wide ranging. Thus the evaluation of (10) is likewise done for a diverse range of conditions.

5. L_N/Q for 2 Cities

The dependence of L_N/Q on NO_x was illustrated in *Kleinman et al.* [2001] using photochemical calculations based on aircraft data collected in Phoenix, AZ. Figure 2a, reproduced from that study, shows an almost one to one correspondence between L_N/Q and NO_x. The few points not lying on the main sequence are from calculations based on

samples with very high toluene concentrations, likely due to canisters with a leaky valve. For Phoenix we find that high NO_x conditions ($L_N/Q > 0.9$, $P(O_3)$ given by (4)) apply whenever NO_x is greater than 3 ppb. Low NO_x conditions ($L_N/Q < 0.1$, $P(O_3)$ given by (3)) apply whenever NO_x is lower than about 0.3 ppb.

Figure 2b shows a similar plot of L_N/Q vs. NO_x based on aircraft data collected in Houston, TX. Compared with Phoenix, points are shifted to higher NO_x concentration and there is much more scatter. Of the 5 cities that we have studied, Phoenix and Houston represent 2 extremes. In Phoenix, emissions are dominated by mobile sources yielding a narrow range of ambient NO_x to VOC_R ratios. Houston has a more diverse emissions mixture which on average has a NO_x to VOC_R ratio lower than Phoenix. It is still possible to pick out low and high NO_x conditions in Houston based on NO_x concentration, but quantifying the middle range without a detailed calculation or evaluation of (17) – (19) is problematic.

6. Formula Evaluation

Our objective is to test Equation (10). To do so we will compare values of $P(O_3)$ obtained from a CSS box model with those determined from Q^{C1} [NO or NO_2]^{C2} $(VOC_R)^{C3}$, which within a proportionality constant is equal to $P(O_3)$. Q, [NO or NO_2], and VOC_R are determined from rate constant data and observed concentrations, temperature, and UV solar irradiance. A CSS model calculation, based on the same set of observations, is used to calculate L_N/Q which according to (7-9) determines the exponents C1, C2, and C3 appearing in the power law formula for $P(O_3)$. Our evaluation procedure has to take into account the fact that the proportionality constant, K, in (10) depends on L_N/Q and is not a priori determined.

There are 624 CSS data points available for evaluating (10). City by city totals are 92 from Nashville, 79 from NYC, 123 from Phoenix, 138 from Philadelphia, and 192 from Houston. The data set has been divided into 10 subsets in which L_N/Q varies from 0

to 1 in increments of 0.1. The first subset has L_N/Q between 0 and 0.1, the second between 0.1 and 0.2, etc. As a way of stressing the wide range of conditions represented by the calculations, Table 2 shows the minimum, maximum, and median values of $P(O_3)$ for each L_N/Q subset. Lowest values of $P(O_3)$ have L_N/Q in the range 0-0.1 or 0.9-1.0. The lowest decile of L_N/Q contains samples from very clean air, some of which are calculated to have net O_3 destruction after taking into account O_3 chemical loss processes. The highest L_N/Q decile is mainly composed of samples with high NO_x concentrations, in a few instances more than 50 ppb due to power plant plumes. Photochemical production of O_3 is then greatly diminished by reaction of OH with NO_2. Highest values of $P(O_3)$ are from Houston [*Kleinman et al.*, 2002a] and occur at intermediate values of L_N/Q.

For each subset we plot the CSS value of $P(O_3)$ versus Q^{C1} [NO or NO_2]^{C2} $(VOC_R)^{C3}$. Values of C1, C2, and C3 are determined for the midpoint of each L_N/Q interval according to (7-9). As K is not predicted beforehand, we are testing the functional form of (10), not whether it gets the right value for $P(O_3)$. By construction, K has zero first derivatives and is therefore nearly constant over a small range of L_N/Q. Thus we expect that our test plots will yield straight lines. Figure 3 contains 10 panels which show plots of $P(O_3)$ versus Q^{C1} [NO or NO_2]^{C2} $(VOC_R)^{C3}$ for the 10 data subsets defined by L_N/Q. Each panel has 2 plots, one using NO and one using NO_2 as the independent variable. A linear least squares regression has been done for the 20 plots yielding the straight line fits shown in Fig. 3. Results are summarized in Table 2. Correlation coefficients show that for L_N/Q between 0 and 0.6, (10) is best evaluated using NO; for $L_N/Q \geq 0.7$, NO_2 gives a higher r^2. NO and NO_2 perform equally well for L_N/Q in the range 0.6 to 0.7, as the power law expression predicts zero dependence on NO_x at $L_N/Q=2/3$. Allowing for a switch in independent variable from NO to NO_2 at $L_N/Q=0.7$, the power law formula explains 96 to 99% of the variability in $P(O_3)$. Figure

3 shows that our formula does equally well in reproducing the CSS results for both low and high $P(O_3)$ cases.

Equation (10) was derived using relative sensitivities that give the response of $P(O_3)$ to small changes in NO_x, VOC_R, and Q (7-9). For that reason neither (10) or the plots in Fig. 3 should be used to infer the change in $P(O_3)$ due to large changes in O_3 precursors. Large changes are accompanied by a change in L_N/Q which in essence shifts the system from one set of exponents to another. In numerical experiments we have found the relative sensitivities to be applicable when there are 10% changes in precursors. Larger changes were not tried.

6.1 Extensions

In order to apply (10) one first needs to know L_N/Q. In our case, L_N/Q comes from model output. To a reasonable degree of accuracy, L_N/Q can also be determined from the approximate equations, (17) – (19). Low and high NO_x regimes where $P(O_3)$ is given by (3) and (4), respectively, can usually be identified based only on NO_x concentration. It is possible that with a limited set of measurements 2 or 3 additional L_N/Q categories could be distinguished. This will certainly be easier to do in some locations than in others; compare Phoenix with Houston.

An alternate way of using (10) would be to rely totally on concentration fields generated from a time dependent Lagrangian or Eulerian model. In that case model output would be used to determine NO_x concentration, VOC reactivity, radical production rate, L_N/Q and $P(O_3)$. An example of such a calculation is given by *Spirig et al.* [2002], in which a Lagrangian model was used to calculate $P(O_3)$ sensitivities over a larger spatial domain than that covered by surface measurement sites. With model output in hand, Equation (10) could be used to categorize variations in $P(O_3)$.

7. Conclusions

An analytic formula has been derived that shows that $P(O_3)$ is proportional to $Q^{C1}[NO_x]^{C2}(VOC_R)^{C3}$, where Q is a radical production rate, $[NO_x]$ is either NO or NO_2 concentration, VOC_R is the sum total reactivity of all VOCs with OH radical, and the exponents, C1, C2, and C3 depend only on L_N/Q, which is the fraction of free radicals removed from the atmosphere by reacting with NO_x. Aside from an undetermined proportionality constant, the power law formula reduces to previously derived results for low and high NO_x conditions. It shows that there is a smooth transition in behavior in the mid-NO_x regime between the two limiting cases.

Using data that we have accumulated in 5 field programs conducted in Nashville, New York City, Phoenix, Philadelphia, and Houston, we show that the power law expression in most cases quantitatively reproduces the functional form of $P(O_3)$ as predicted from a constrained steady state (CSS) box model. Our analysis shows that the power law formula and CSS model are consistent but does not establish the accuracy of our formula because that depends on the accuracy and completeness of the CSS calculations. One item that we know is missing is the effect of NO_x reservoir compounds such as PAN. We therefore expect our formula to be most accurate in conditions where PAN concentrations are low or where PAN is in steady state. Even with a sizable PAN concentration we have reason to believe that $P(O_3)$ sensitivities and the power law formula for $P(O_3)$ are qualitatively useful.

Just as low and high NO_x formulas for $P(O_3)$ have been helpful in analyzing field data and understanding day to day and place to place variability in $P(O_3)$, we hope that the power law expression will have a similar function in the more general mid-NO_x range. Application of the general formula to mid-NO_x conditions is, however, more demanding as a value for L_N/Q is needed. An analysis of model output, rather than measurements, would naturally yield L_N/Q. The power law formula could then be applied to the task of understanding how and why $P(O_3)$ varies over the model domain.

Appendix: dlnP(O_3)/dlnQ

Equation (9) for dlnP(O_3)/dlnQ is derived following the same procedures used to determine dlnP(O_3)/dln[NO] and dlnP(O_3)/dln(VOC_R) [*Kleinman, et al.*, 1997]. Ozone production occurs by the reaction of peroxy radicals with NO in R5-R7 followed by photolysis of NO_2, in R8. The O_3 production rate can therefore be expressed as

$$P(O_3) = k_t ([HO_2]+[RO_2]) [NO] \quad (A1)$$

where k_t is a weighted average rate constant for R5 – R7 (see *Kleinman et al.*, 1997). The production rate for radicals (odd hydrogen = OH + HO_2 + RO_2), Q, must equal the sum of radical sinks:

$$Q = 2k_{12} [HO_2]^2 + 2k_{13} [HO_2][RO_2] + L_R + L_N \quad (A2)$$

The first 2 terms in (A2) represent loss of radicals due to production of H_2O_2 and organic peroxides. L_R represents all other radical - radical reactions including OH+HO_2 and RO_2 + R'O_2, as well as first order loss processes. L_N includes all radical loss reactions between free radicals and NO or NO_2 (R9-R11). Equation (A2) can be re-arranged to give the total peroxy radical concentration:

$$[HO_2]+[RO_2] = (Q - L_R - L_N)^{1/2} / (2 k_{eff})^{1/2} \quad (A3)$$

where k_{eff} is an effective rate constant for peroxide formation,

$$k_{eff} = k_{12}(1-\alpha')^2 + k_{13}(1-\alpha') \alpha' \quad (A4)$$

$$\alpha' = [RO_2]/([HO_2] + [RO_2])$$

Substituting (A3) into (A1) gives

$$P(O_3) = k_t/(2 k_{eff})^{1/2} (Q - L_R - L_N)^{1/2} [NO] \quad (A5)$$

At this point we make our first significant approximation by setting L_R equal to zero. A relative sensitivity of $P(O_3)$ to Q (i.e., dlnP/dlnQ = (Q/P)dP/dQ) is obtained by differentiating (A5) with respect to Q.

$$\mathrm{dlnP(O_3)/dlnQ} = 1/2Q\ (1 - dL_N/dQ)/(Q - L_N) \qquad \text{(A6)}$$

A more serious approximation is now made by ignoring organic nitrate and PAN formation leaving just HNO_3 production to contribute to L_N:

$$L_N \approx P(HNO_3) = k_9[OH][NO_2] \qquad \text{(A7)}$$

The OH concentration appearing in (A7) is removed from our equations in favor of $P(O_3)$ by expressing $P(O_3)$ as

$$P(O_3) \approx \sum Y_i\, k_i\, [OH]\, [VOC_i] \qquad \text{(A8)}$$

Equation (A8) is an approximation which becomes exact if all peroxy radicals formed from reactions of OH with VOCs go on to react with NO forming O_3. *Tonnesen and Dennis* [2000a, Fig. 1d] present results that address part of (A8), namely the fraction of HO_2 that reacts with NO. Except at low NO_x concentration that fraction is 80 – 100%, sufficiently close to unity that (A8) is a useable approximation. However, at low NO_x the primary fate of a peroxy radical is to form peroxides. That (A8) is no longer valid at low NO_x has very little effect on our final results as under that condition L_N and its derivatives are close to zero.

Substituting (A7) into (A8) yields,

$$L_N = (k_9[NO_2] / \sum Y_i\, k_i\, [VOC_i]\ )\ P(O_3) \qquad \text{(A9)}$$

Equation (A9) is differentiated with respect to Q, yielding

$$dL_N/dQ = (L_N/Q)\, dlnP(O_3)/dlnQ \quad \text{(A10)}$$

which is substituted into (A6). After collecting like terms we obtain

$$dlnP(O_3)/dlnQ = (1/2)/(1 - 1/2\, L_N/Q) \quad \text{(A11)}$$

Acknowledgement

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Table 1 Dependence of $P(O_3)$ on Q, NO_x, and VOC_R

L_N/Q	Exponents for $P(O_3) = K\, Q^{C1}\, [NO_x]^{C2}\, (VOC_R)^{C3}$		
	C1	C2	C3
0.0	0.5	1.0	0.0
0.5	0.67	0.33	0.33
0.67	0.75	0.0	0.5
1.0	1.0	-1.0	1.0

Table 2. Characteristics of CSS calculated $P(O_3)$ and comparison with $P(O_3)$ power law formula.

L_N/Q	# samples	$P(O_3)$ (ppb h^{-1})			linear least squares[1]: r^2	
		Min.	Max.	Median	NO	NO_2
0.0 – 0.1	136	0.04	12	2.7	0.96	0.87
0.1 – 0.2	92	0.7	20	6.5	0.97	0.84
0.2 – 0.3	61	1.4	71	8.4	0.99	0.86
0.3 – 0.4	35	1.9	22	11	0.97	0.86
0.4 – 0.5	25	2.0	155	13	0.99	0.97
0.5 – 0.6	29	1.9	46	15	0.97	0.92
0.6 – 0.7	16	3.0	46	17	0.97	0.97
0.7 – 0.8	32	2.9	67	22	0.97	0.99
0.8 – 0.9	34	2.5	32	15	0.89	0.98
0.9 – 1.0	164	0.6	61	7.4	0.91	0.99

[1] From correlation of CSS $P(O_3)$ versus Q^{C1} [NO or NO_2]^{C2} $(VOC_R)^{C3}$.

Figure Captions

1. Relative sensitivity of $P(O_3)$ to radical production rate, Q; NO_x concentration; and VOC reactivity, VOC_R, as a function of the fraction of radicals removed by reaction with NO_x, L_N/Q. Curves are from (7-9).

2. The dependence of L_N/Q on log of NO_x concentration in (a) Phoenix and (b) Houston. Each point represents a sampling location where a CSS calculation was done yielding a value for L_N/Q.

3. Calculated values of $P(O_3)$ from a CSS box model versus power law formula for $P(O_3)$. The CSS calculations have been divided into 10 subsets in which L_N/Q varies from 0 to 1 in increments of 0.1 as indicated on the 10 panels. In each panel the power law expression for $P(O_3)$, Q^{C1} [NO or NO_2]^{C2} $(VOC_R)^{C3}$, is evaluated using NO (shaded symbols) and NO_2 (open symbols). C1, C2, and C3, are determined from (7-9) using the midpoint of the L_N/Q range. Straight lines are linear least squares fit to data points. The squared correlation coefficients, r^2, are given in Table 2. Symbols shown in the first panel identify points as based on data collected in Nashville, TN; New York City, NY; Phoenix, AZ; Philadelphia, PA, or Houston, TX.

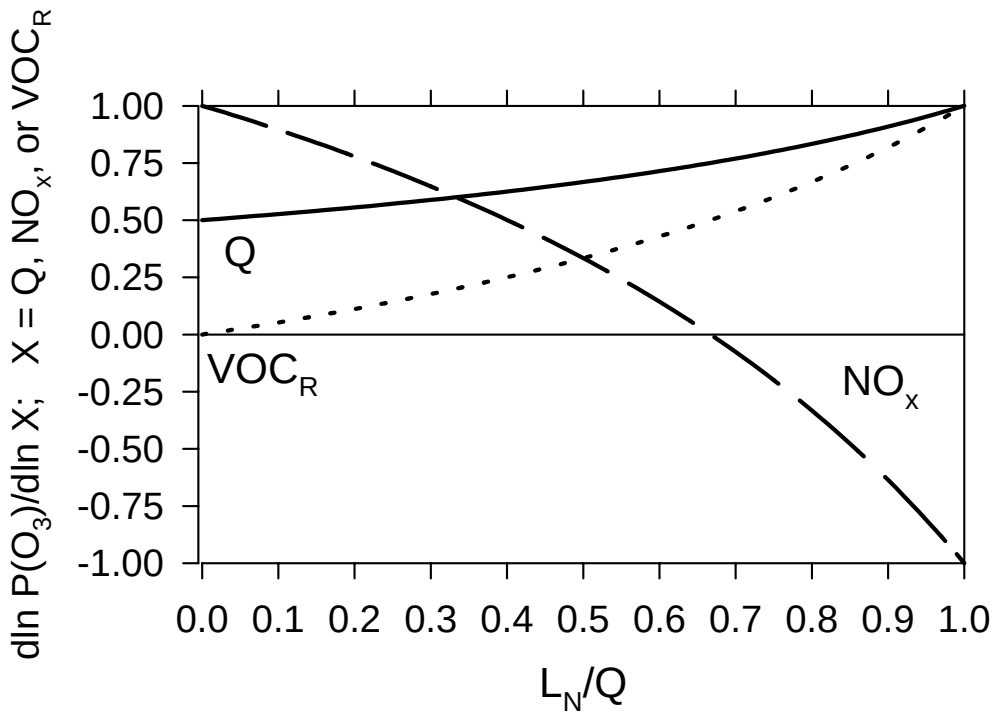

dln P(O3)/dln X; X = Q, NOx, or VOCR
1.00
0.75
0.50
0.25
0.00
-0.25
-0.50
-0.75
-1.00
Q
VOCR
NOx
0.0 0.1 0.2 0.3 0.4 0.5 0.6 0.7 0.8 0.9 1.0
LN/Q

Figure 1

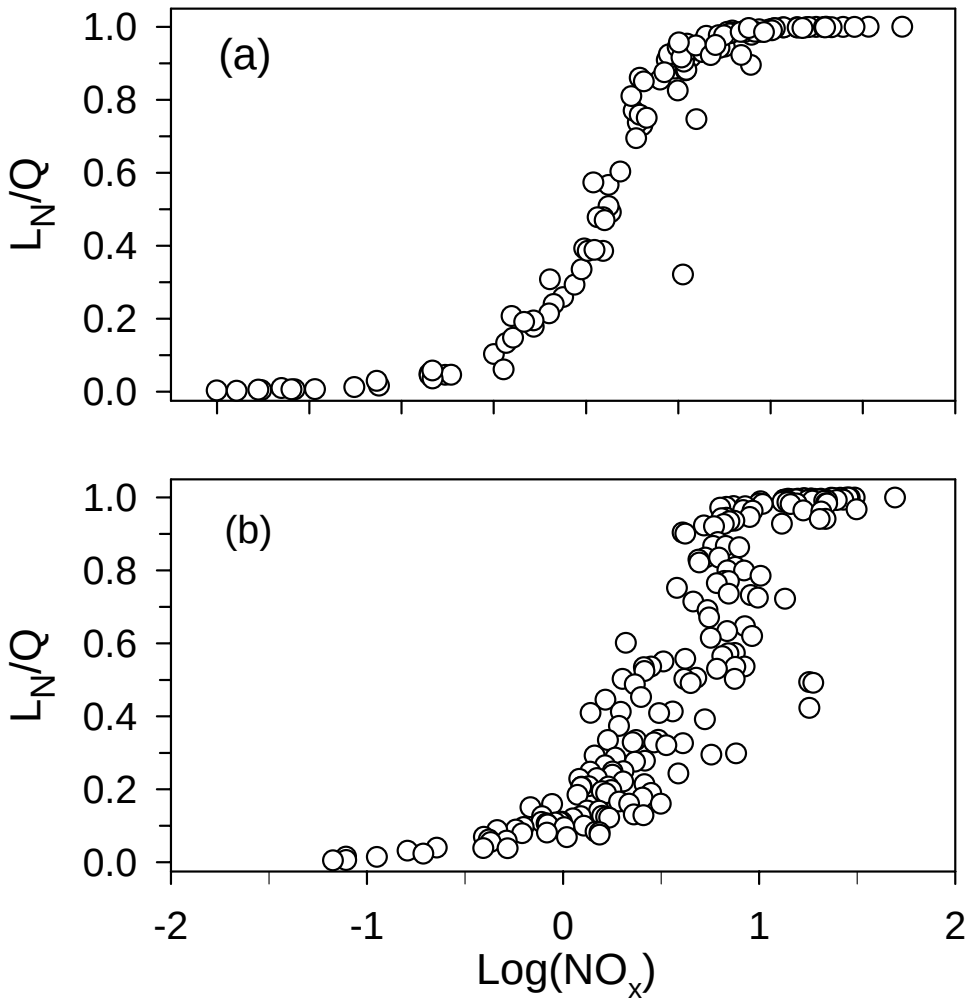


Figure 2

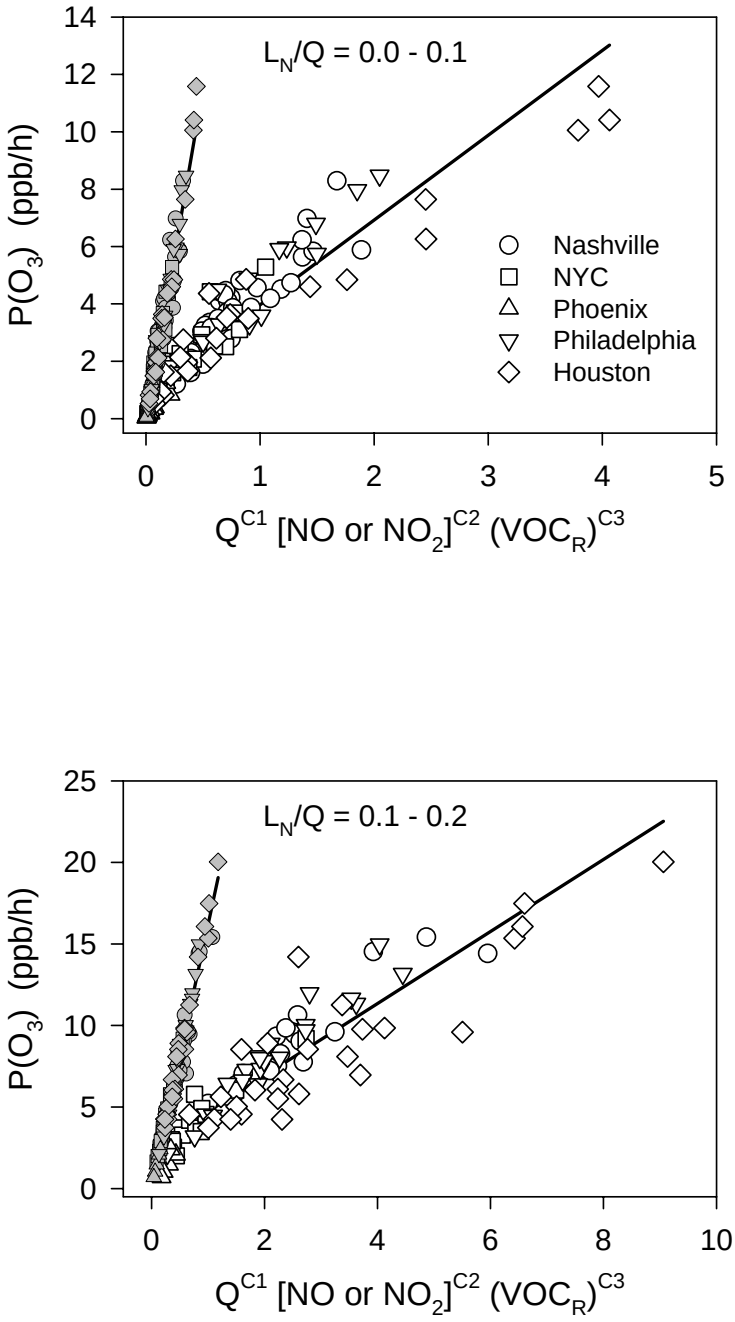
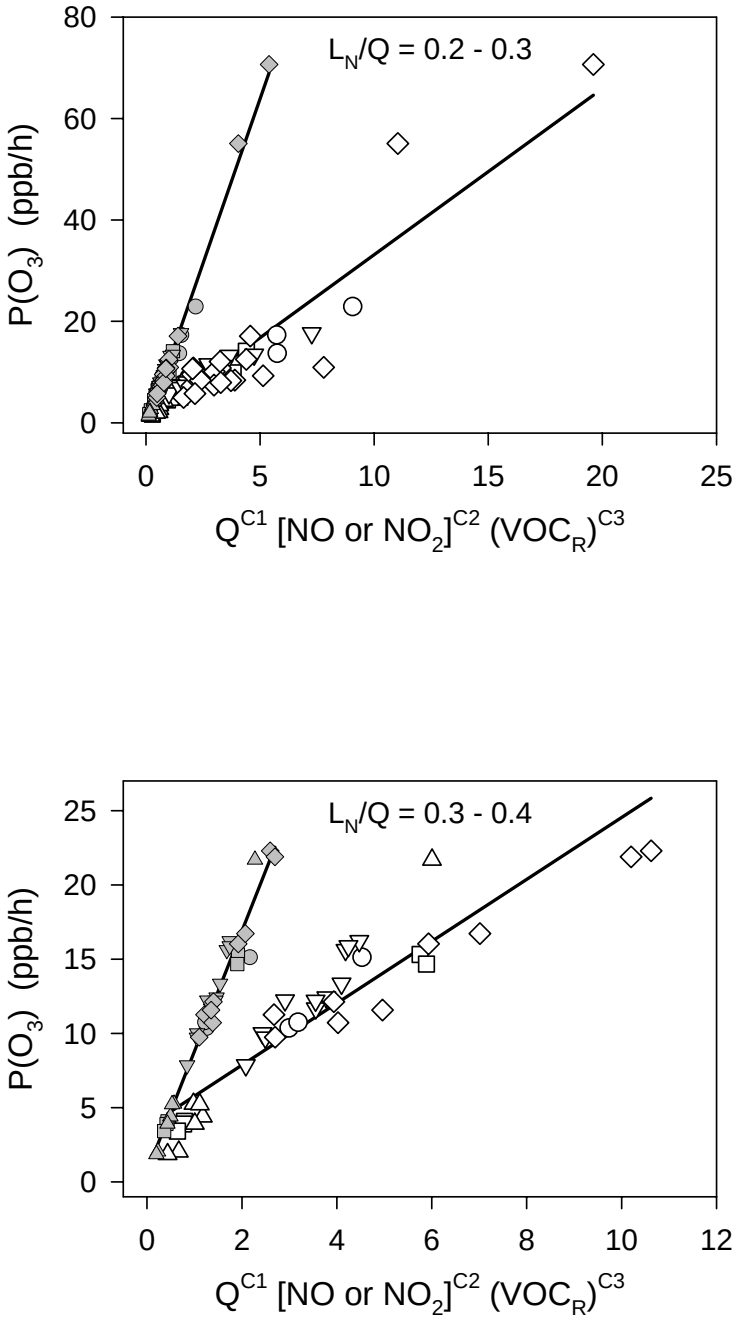

L_N/Q = 0.0 - 0.1
$P(O_3)$ (ppb/h)
Q^{C1} [NO or NO_2]^{C2} (VOC_R)^{C3}
Nashville
NYC
Phoenix
Philadelphia
Houston
L_N/Q = 0.1 - 0.2
$P(O_3)$ (ppb/h)
Q^{C1} [NO or NO_2]^{C2} (VOC_R)^{C3}

Figure 3



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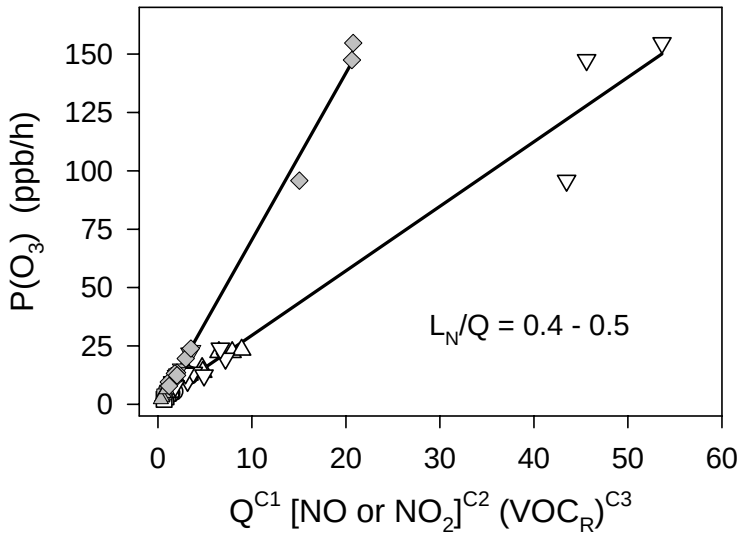


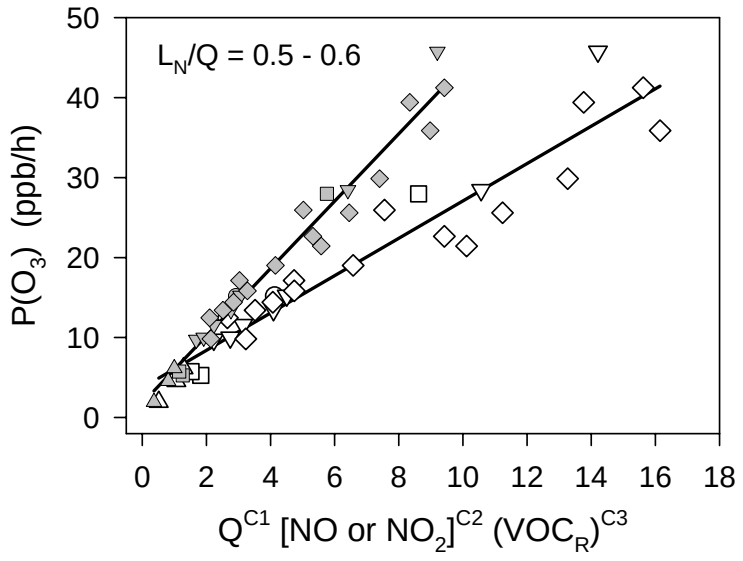
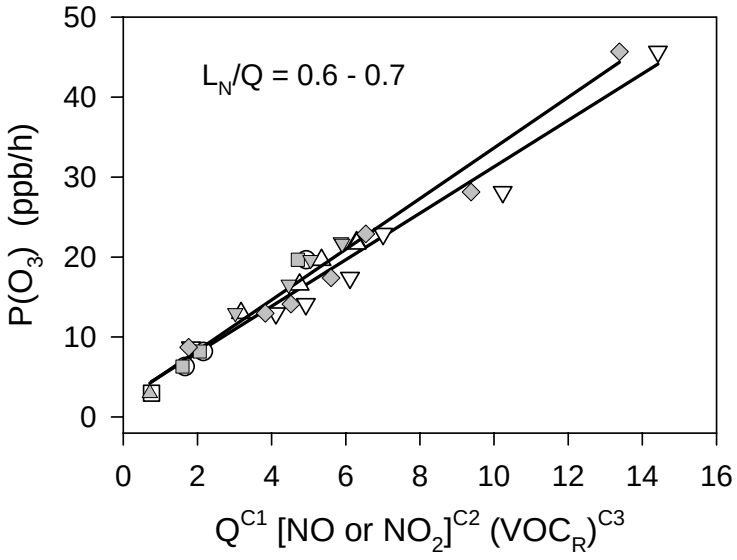
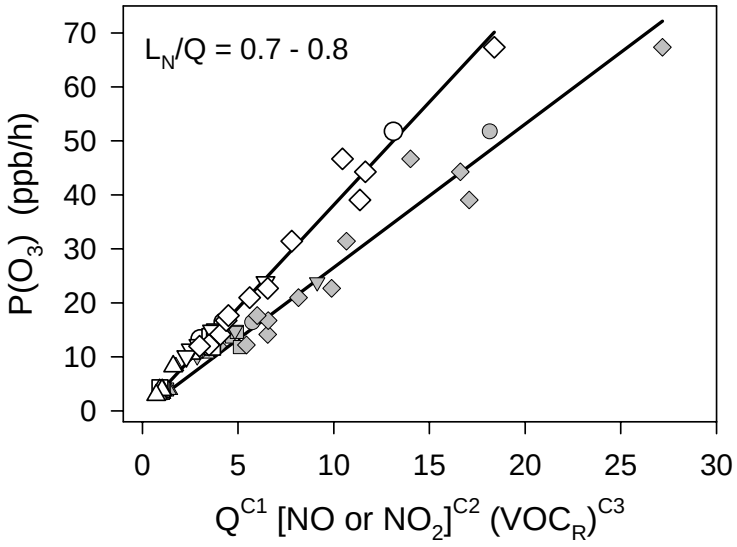


Figure 3, continued





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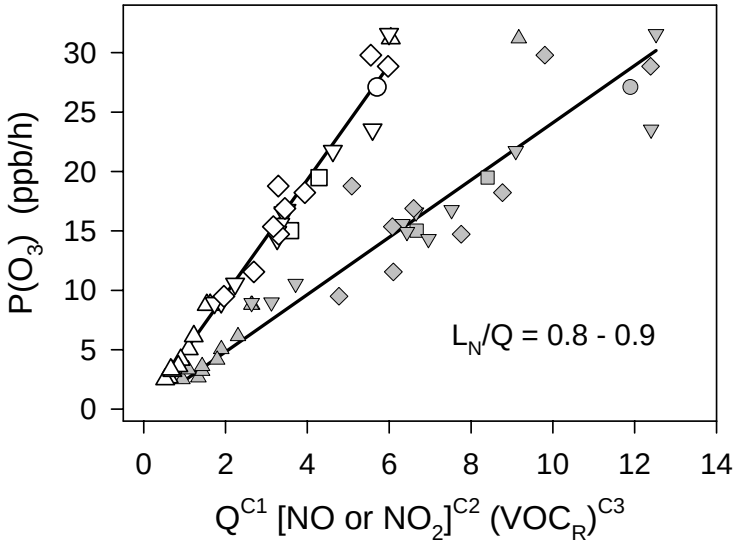

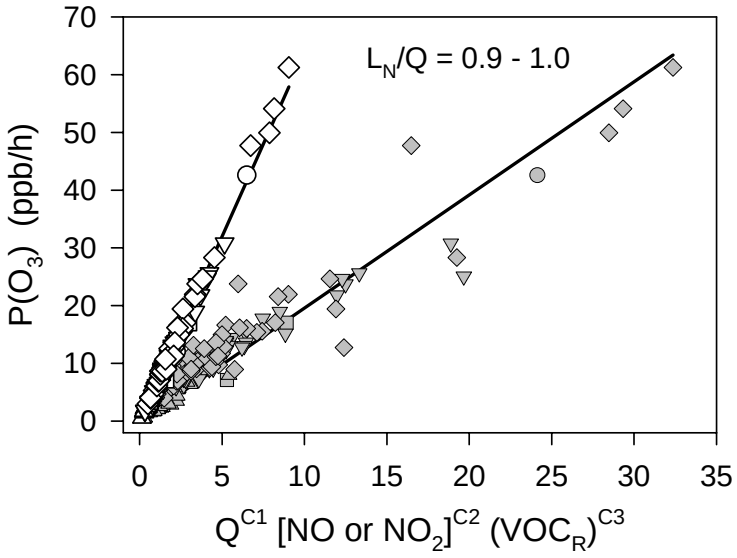

Figure 3, continued